

SYNTHESIS OF OLIGOMERIC POLY(SILARYLENE-SILOXANE-ACETYLENE)'S  
AND THEIR CONVERSION TO HIGH TEMPERATURE PLASTICS,  
ELASTOMERS, AND COATINGS

BACKGROUND OF THE INVENTION

1. Field of the Invention

[0001] This invention relates to silarylene-siloxane prepolymers, silarylene-siloxane-acetylene precursors, silarylene-siloxane-acetylene networked polymers, silicon-containing ceramic compositions, and processes for making the same.

2. Description of the Related Art

[0002] Functionalized organosilicon polymers represent a diverse class of materials that have found widespread use in a variety of applications, ranging from contact lenses, to medical implant devices to high-temperature materials for automotive and aircraft parts. Polysiloxanes or silicone polymers have been mostly studied and are of the greatest commercial importance. A major limitation of organosiloxane polymers is their limited long-term thermal stability at temperatures in excess of 200°C caused primarily by ionic degradation reactions and formation of stable cyclic products such as six- and/or eight-membered siloxane rings. To improve long term thermal stability at high temperatures while retaining low temperature flexibility, various investigations have been directed toward the incorporation of aromatic units into the backbone or into the pendant groups to prevent formation of the cyclic products. However, some synthesis routes for these aromatic organosilicon polymers still result in degradative side reactions that limit molecular weight.

[0003] Silarylene-siloxane polymers,  $-\text{[Si(R}_2\text{)-Ar-}\{\text{Si(R}_2\text{)-O-}\}_z\text{]-}$ , in which a fraction of the total content of siloxy oxygens in the polysiloxane chain are replaced by aromatic units, are the most studied class of aromatic containing polymers. Four primary routes for the syntheses of silarylene-siloxane polymers in which one of the aforementioned difunctional monomers is allowed to react with an arylenedisilanol have been disclosed: chlorosilane (Lai et al., *J. Polym. Sci. Polym. Chem. Ed.*, 1982, 20, 2277), acetoxysilane (Rosenberg et al., *Polym. Preprints*, 1978, 19(2), 625), aminosilane (Burks et al., *J.*

*Polym. Sci. Polym. Chem. Ed.*, 1973, 11, 319), and ureidosilane (Dvornic et al., *W. J. Polym. Sci. Polym. Chem. Ed.*, 1982, 20, 951) polycondensation methods. The occurrence of degradative side reactions between acidic by-products and the growing polymer chain has been shown to limit molecular weights and disrupt the truly alternating nature of the polymer structures prepared via the former two methods. The aminosilane method produces a weakly basic by-product. The ureidosilane route has been reported to be the method of choice for producing high-molecular weight linear silarylene-siloxane polymers with exactly alternating structures. The success of the ureidosilane method is due to the formation of an unreactive polycondensation by-product that reduces the occurrence of degradative side reactions.

**[0004]** Zhang et al. disclosed a process of making high-molecular-weight silphenylene-siloxane homopolymers and copolymers via dehydrocoupling polymerization using Wilkinson's catalyst ( $(\text{Ph}_3\text{P}_3)\text{RhCl}$ ). (*Macromolecules* 2000, 33, 3508). The compositions were made using a 1:1 molar ratio of the two monomers and did not have specific terminating groups. U.S. Patent 5,874,514 to Keller et al. discloses linear inorganic-organic hybrid polymers having repeat units that contain at least one alkynyl group and at least one siloxanyl group within the backbone. U.S. Patent 5,563,181 to Keller et al. discloses inorganic-organic hybrid thermoset polymers that are formed from linear inorganic-organic hybrid polymers having repeat units that contain at least one alkynyl group for cross-linking purposes and at least one siloxanyl group. U.S. Patents 5,346,980 and 5,578,380 to Babu disclose crosslinkable copolymers suitable for use as elevated temperature pressure-sensitive adhesives comprising randomly arranged silarylene units and siloxane units. U.S. SIR H1612 to Rhein et al. discloses a method of making silarylene-siloxane polymers.

**[0005]** There is a need for oxidatively stable crosslinked silarylene-siloxane networked polymers. The networked polymers should be made by a process that avoids degradative side reactions that limit molecular weight and reduce performance at high temperatures. Such networked polymers are useful for high temperature plastics, elastomers, adhesives, and coatings. There is also a need for a hydroxy-terminated prepolymer needed to form the networked polymers. There is a further need for a

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ceramic composition formed from such networked polymers. Such materials would be useful in reinforced composites.

## SUMMARY OF THE INVENTION

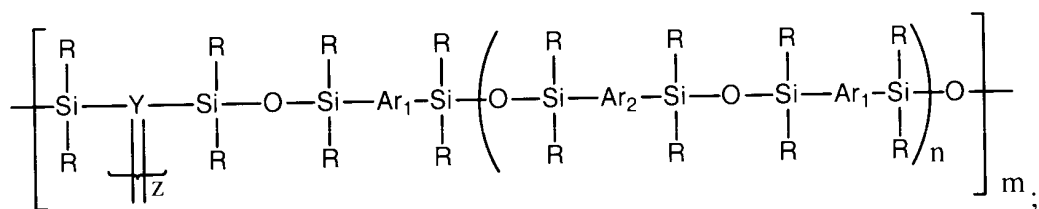
[0006] An object of the invention is to provide oxidatively stable silarylene-siloxane-acetylene networked polymers suitable for use as high temperature plastics, elastomers, adhesives, and coatings, capable of retaining their elastomeric properties up to 250°C.

[0007] A further object of the invention is to provide prepolymers and precursors that are needed to form the networked polymers of the invention.

[0008] A further object of the invention is to provide ceramic compositions, made by pyrolysis of the networked polymers, that are useful in reinforced composites.

[0009] A further object of the invention is to provide processes for making the networked polymers, prepolymers, and precursors of the invention, while avoiding degradative side reactions.

[0010] These and other objects of the invention are accomplished by a networked polymer comprising the formula:



wherein  $n \geq 1$ ;

wherein  $n$  is an average value obtained by averaging all repeating units of the networked polymer;

wherein  $m \geq 1$ ;

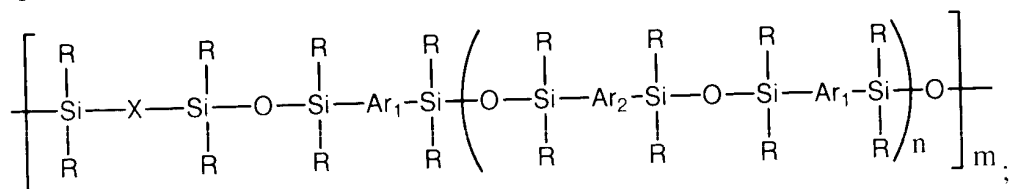
wherein  $Y$  is a divalent group containing one or more acetylenic groups, one or more ethenyl crosslinks, or both;

wherein  $z$  is the crosslink density;

wherein  $\text{Ar}_1$  and  $\text{Ar}_2$  are independently selected aromatic groups; and

wherein each  $R$  is independently selected from the group consisting of alkyl, aryl, alkylaryl, haloalkyl, haloaryl, and combinations thereof.

[0011] A further embodiment of the invention is a precursor comprising the formula:



wherein  $n \geq 1$ ;

wherein  $n$  is an average value obtained by averaging all repeating units of the precursor;

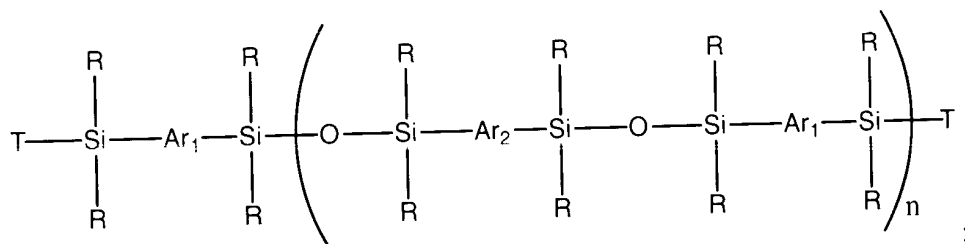
wherein  $m \geq 1$ ;

wherein  $X$  is a divalent group containing one or more acetylenic groups;

wherein  $\text{Ar}_1$  and  $\text{Ar}_2$  are independently selected aromatic groups; and

wherein each  $R$  is independently selected from the group consisting of alkyl, aryl, alkylaryl, haloalkyl, haloaryl, and combinations thereof.

[0012] A further embodiment of the invention is a prepolymer comprising the formula



wherein  $n \geq 1$ ;

wherein  $T$  is either  $-\text{H}$  or  $-\text{OH}$  and both  $T$ 's are the same;

wherein  $\text{Ar}_1$  and  $\text{Ar}_2$  are independently selected aromatic groups; and

wherein each  $R$  is independently selected from the group consisting of alkyl, aryl, alkylaryl, haloalkyl, haloaryl, and combinations thereof.

[0013] A further embodiment of the invention is a ceramic composition formed by pyrolysis of the above networked polymer.

[0014] A further embodiment of the invention is a process for making the prepolymer, being a hydroxy-terminated prepolymer or a hydride-terminated prepolymer,

comprising the step of reacting one or more arylenedisilanol with one or more bissilarylenes.

[0015] A further embodiment of the invention is a process for making the precursor comprising the step of reacting a hydroxy-terminated prepolymer with a bis(dimethylaminosilyl)alkyne.

[0016] A further embodiment of the invention is a process for making the networked polymer comprising the step of crosslinking the precursor.

#### BRIEF DESCRIPTION OF THE DRAWINGS

[0017] Figure 1 shows a general process for making the hydroxy-terminated prepolymer, the precursor, the networked polymer, and the ceramic composition.

[0018] Figure 2 shows a process for making a hydroxy-terminated prepolymer, precursor, networked polymer, and ceramic composition using specific reactants.

[0019] Figure 3 shows a multi-step process for making a hydroxy-terminated prepolymer via a hydride-terminated prepolymer.

#### LIST OF REFERENCE NUMBERS

[0020]	100	arylenedisilanol
[0021]	102	bissilarylene
[0022]	104	hydroxy-terminated prepolymer
[0023]	106	bis(dimethylaminosilyl)alkyne
[0024]	108	precursor
[0025]	110	networked polymer
[0026]	112	ceramic composition
[0027]	200	1,4-bis(hydroxydimethylsilyl)benzene
[0028]	202	1,4-bis(dimethylsilyl)benzene
[0029]	204	hydroxy-terminated prepolymer
[0030]	206	1,4-bis(dimethylaminodimethylsilyl)butadiyne
[0031]	208	precursor
[0032]	210	networked polymer
[0033]	212	ceramic composition

[0034]	300	arylenedisilanol
[0035]	302	bissilarylene
[0036]	303	hydride-terminated prepolymer
[0037]	304	hydroxy-terminated prepolymer
[0038]	305	arylenedisilanol

## DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

[0039] The synthesis of the networked polymer 110 of the invention is performed in three reaction steps. Each step is a novel process and produces a novel composition. One additional step can be performed to produce the ceramic composition 112. Figure 1 shows the general scheme for the reaction. First, one or more arylenedisilanol 100 and one or more bissilarylenes 102 are reacted to form a hydroxy-terminated prepolymer 104. The hydroxy-terminated prepolymer 104 is then reacted with a bis(dimethylaminosilyl)alkyne 106 to form a precursor 108. The precursor 108 is then crosslinked to form a networked polymer 110.

### 1. Formation of the prepolymer

[0040] In the first step one or more arylenedisilanol ( $\text{HORRSi}-\text{Ar}_1-\text{SiRR'OH}$ ) 100 are reacted with one or more bissilarylenes ( $\text{HRRSi}-\text{Ar}_2-\text{SiRR'H}$ ) 102 to make a prepolymer. Preferably, only one arylenedisilanol 100 compound and only one bissilarylene 102 compound are used. As shown in Figure 1, the bissilarylene 102 must have one or more hydride groups on each silicon atom, and the term bissilarylene in this application refers only to such compounds. The  $\text{Ar}_1$  and  $\text{Ar}_2$  are independently selected aromatic groups. The aromatic group can be any group containing one or more linear or fused aromatic rings, with or without intervening functional groups, wherein both silicon atoms are bonded directly to the same or different aromatic rings. Phenylene is a preferred aromatic group. 1,4-Phenylene is a more preferred aromatic group. The general structures are shown using only  $\text{Ar}_1$  and  $\text{Ar}_2$  to illustrate the alternating addition of arylenedisilanol 100 and bissilarylene 102. When multiple arylenedisilanol 100 and bissilarylenes 102 are used,  $\text{Ar}_1$  and  $\text{Ar}_2$  may each represent more than one different aromatic group. Each R represents a functional group selected from the group consisting

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of alkyl, aryl, alkylaryl, haloalkyl, haloaryl, and combinations thereof. Methyl is a preferred R group. More preferably, every R group is methyl. Each aromatic group and R group can be selected independently from the others when used in any process step or composition of the invention. The aromatic groups and R groups do not directly participate in the reactions of the invention. Thus the process should not be significantly affected regardless of the specific aromatic groups and R groups chosen.

[0041] The particular value for  $n$  and the particular choices for the R groups and aromatic groups may be selected according to the particular properties desired for the precursor 108, networked polymer 110, and ceramic composition 112. Increasing the relative number of silarylene and siloxane units (by increasing  $n$ ) increases the chain flexibility. Using longer chain R groups increases the solubility of the precursor 108 in organic solvents. Longer chain R groups also increase the hydrophobicity and decrease the thermo-oxidative stability of the networked polymer 110. Aryl-containing R groups increase the stiffness and slightly increase the thermo-oxidative stability. Larger aromatic groups (Ar) add stiffness.

[0042] The integer  $n$  must be greater than or equal to 0. A novel compound is formed when  $n$  is greater than or equal to 1. When  $n=0$ , the reaction is unnecessary and an unreacted arylenedisilanol 100 is used as the hydroxy-terminated prepolymer 104. An unreacted bissilarylene 102 is not suitable for the next reaction step. It should be noted that the repeating unit of the prepolymer contains both an arylenedisilanol 100 unit and a bissilarylene 102. Therefore the prepolymer contains  $2n+1$  arylenedisilanol 100 and bissilarylene 102 units.

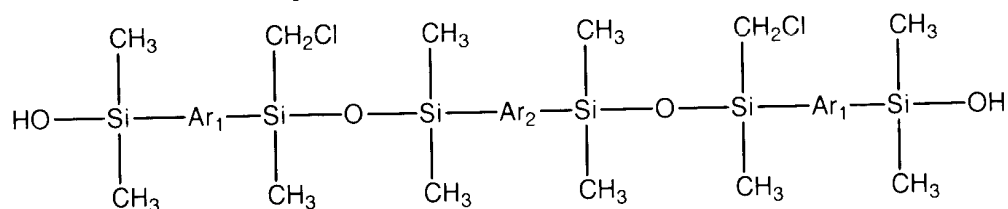
[0043] The reaction of excess arylenedisilanol 100 with bissilarylene 102 produces an oligomeric hydroxy-terminated prepolymer 104. The oligomeric chain is formed by the dehydrogenative coupling of the two reactants. When arylenedisilanol 100 reacts with bissilarylene 102,  $H_2$  is lost, creating a siloxane linkage between the two reactants. The chain grows longer by alternating addition of arylenedisilanol 100 and bissilarylene 102 and loss of  $H_2$  with each addition.

[0044] The length of the hydroxy-terminated prepolymer 104 chain can be controlled by adjusting the molar ratio of the reactants. The arylenedisilanol 100 must be present in an excess amount to ensure hydroxy termination on both ends of the chain. For example,

a 3:2 molar ratio of arylenedisilanol 100 and bissilarylene 102 produces an average chain length of 5 units and an average n of 2. The average hydroxy-terminated prepolymer 104 contains 3 units of arylenedisilanol 100 and 2 units of bissilarylene 102 in an alternating structure, with hydroxy termination on both ends of the chain. The terminal hydroxy groups cannot react with the arylenedisilanol 100. Once all the bissilarylene 102 has been consumed and terminated with arylenedisilanol 100, reaction will cease.

[0045] A novel composition is also produced when the bissilarylene 302 is present in an excess amount. This scheme is illustrated in Figure 3. Excess bissilarylene 302 results in a hydride-terminated prepolymer 303. The length of the hydride-terminated prepolymer 303 is controlled in the same way as that of the hydroxy-terminated prepolymer 104. The hydride-terminated prepolymer 303 is within the scope of the invention. However, it must be converted to a hydroxy-terminated prepolymer 304 in order to make the precursor 108. This conversion can be done by any means known in the art that results in a hydroxy-terminated prepolymer 104 with the structure shown in Figure 1. This can be done by a reaction that converts the terminal hydrides to hydroxies without affecting any other atoms of the hydride-terminated prepolymer 303. It can also be done by increasing the chain length of the prepolymer. Figure 3 illustrates such a scheme. The hydride-terminated prepolymer 303 is reacted with additional arylenedisilanol 305, which can be different from the first arylenedisilanol 300. The resulting hydroxy-terminated prepolymer 304 can then be used in the next reaction step.

[0046] The arylenedisilanol 100 and bissilarylene 102 may have an asymmetrical arrangement of R functional groups. However, the reaction may first occur at either of the hydroxy or hydride groups in a particular molecule. Therefore some repeating units may have a mirror-image arrangement of R groups relative to the other repeating units. Each possible arrangement is an enantiomer of the same prepolymer. This is illustrated in the structure below of a specific hydroxy-terminated prepolymer 104 with n=1.





**[0047]** This hydroxy-terminated prepolymer 104 contains two units of a arylenedisilanol 100 that contains one chloromethyl group. In the left-most unit the chloromethyl group is to the right of the  $Ar_1$  group and in the right-most unit the chloromethyl group is to the left of the  $Ar_1$  group. Two more enantiomers would have the chloromethyl groups on the same side of the  $Ar_1$  groups, either to the left or the right. The longer the prepolymer chain, the more enantiomers are possible. The general structures of the compositions of the invention include all possible enantiomers of the prepolymer, the precursor 108, and the networked polymer 110.

**[0048]** The advantage of using dehydrogenative coupling is that the formation of molecular hydrogen as a by-product eliminates degradative side reactions that limit molecular weight and disrupt the alternating structure. The completion of the reaction when making a hydroxy-terminated prepolymer 104 can be confirmed by IR analysis of the Si—H stretching band in the bisarylene 102. When this band disappears, the reaction is complete.

**[0049]** Preferably, a catalyst is used in the reaction. Preferably, the reaction is catalyzed with a rhodium-containing compound. Wilkinson's catalyst,  $(Ph_3P_3)RhCl$ , is a preferred rhodium-containing compound. Other possible catalysts include tris(dibenzylideneacetone)dipalladium-chloroform adduct  $(Pd_2(dba)_3-CHCl_3)$ , palladium on activated carbon, dicobalt octacarbonyl  $(Co_2CO_8)$ , and bis(cyclooctadiene)nickel  $Ni(cod)_2$ .

**[0050]** Figure 2 shows preferred reactants for the process. 1,4-bis(hydroxydimethylsilyl)benzene  $(HO(CH_3)_2Si-C_6H_4-Si(CH_3)_2OH)$  200 is reacted with 1,4-bis(dimethylsilyl)benzene  $(H(CH_3)_2Si-C_6H_4-Si(CH_3)_2H)$  202 to produce the hydroxy-terminated prepolymer 204.  $^1H$  NMR and  $^{13}C$  NMR can be used to verify that the hydroxy-terminated prepolymer has a truly alternating structure. In that case, there is a single signal for the methyl protons and methyl carbons.

## 2. Formation of the precursor

**[0051]** In the next reaction step, the hydroxy-terminated prepolymer 104 is reacted with a bis(dimethylaminosilyl)alkyne  $[(CH_3)_2N-RRSi-X-SiRR-N(CH_3)_2]$  106. The X in the bis(dimethylaminosilyl)alkyne 106 is a divalent group containing one or

more acetylenic groups. When there is more than one acetylenic group, the acetylenic groups may or may not be conjugated. The reaction produces the precursor 108. The structure of the precursor 108 alternates between units of the prepolymer 104 and units of the bis(dimethylaminosilyl)alkyne 106. The precursor 108 can be a monomer when  $m=1$ , or can be oligomeric or polymeric when  $m \geq 2$ . Because the synthesis of the hydroxy-terminated prepolymer 104 produces chains of varying lengths,  $n$  in the precursor 108 is an average value greater than or equal to 0, obtained by averaging all repeating units of the precursor 108.

[0052] The reaction can proceed without a catalyst and results in the loss of dimethylamine. Preferably, the bis(dimethylaminosilyl)alkyne 106 is added to the reaction mixture in more than one aliquot.  $^1\text{H}$  NMR can be used to determine whether excess bis(dimethylaminosilyl)alkyne 106 remains, requiring the addition of additional hydroxy-terminated prepolymer 104 to consume the bis(dimethylaminosilyl)alkyne 106.

[0053] Figure 2 shows preferred reactants for the process. 1,4-bis(dimethylaminodimethylsilyl)butadiyne  $[(\text{CH}_3)_2\text{N}-\text{Si}(\text{CH}_3)_2-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-\text{Si}(\text{CH}_3)_2-\text{N}(\text{CH}_3)_2]$  206 is reacted with the hydroxy-terminated prepolymer 204 described above to produce the precursor 208.

### 3. Formation of the networked polymer

[0054] The final reaction step is the crosslinking of the precursor 108 to produce a networked polymer 110.  $Y$  in the networked polymer 110 is a divalent group containing one or more acetylenic groups, one or more ethenyl crosslinks, or both. Within each  $Y$ , a former acetylenic group may be converted to a single bond between ethenyl crosslinks, or a former pair of conjugated acetylenic groups may be converted to a single acetylenic group between ethenyl crosslinks. Other crosslinking schemes are also possible and are known in the art. Crosslinking does not necessarily occur in every repeating unit; thus  $z$  represents the crosslink density. The value  $n$  is an average value greater than or equal to 0, obtained by averaging all repeating units of the networked polymer 110.

[0055] The crosslinking can be performed by heating the precursor 108. The heating conditions can be at least sufficient to initiate crosslinking, but not severe enough to cause degradation of the precursor 108 or the networked polymer 110. Preferably, the

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heating is done from about 100°C to about 500°C. The use of a nitrogen atmosphere allows the heating to be done at the high end of the temperature range. Such high temperatures may cause degradation if it were done under an oxidizing atmosphere.

[0056] The number of crosslinks per unit volume can be controlled by controlling length of the hydroxy-terminated prepolymer 104, which is controlled by the molar ratio of arylenedisilanol 100 to bisilarylene 102 as described above. A shorter hydroxy-terminated prepolymer 104 results in closer spacing of the crosslinkable acetylenic groups in the precursor 108. As the hydroxy-terminated prepolymer 104 gets longer, the spacing of the crosslinkable acetylenic groups increases. The greater the spacing, the lower the number of crosslinks per unit volume. Physical properties typically depend on the degree of crosslinking. Thus, the physical properties can be controlled by controlling the molar ratio of reactants in the first step of the process.

[0057] Figure 2 shows the crosslinking of a preferred precursor 208 described above. The networked polymer 210 illustrates the conversion of a pair of conjugated acetylenic groups to a single acetylenic group between ethenyl crosslinks.

[0058] The process described above can produce networked polymers 110 that are completely insoluble in typical organic solvents. By varying the crosslink density of the networked polymer 110, polymeric materials can be obtained that vary from slightly (elastomeric) to highly (plastic) crosslinked systems. The networked polymers 110 can be oxidatively stable and can retain their elastomeric properties up to about 250°C. When aged at higher temperatures for extended times, they can lose weight and become brittle.

#### 4. Formation of the ceramic composition

[0059] The final composition of the invention is a ceramic composition 112 that can be produced by pyrolysis of the networked polymer 110. The networked polymer 110 should be heated to a temperature at which the thermal degradation commences. Heating can be continued until the weight loss of the material has ceased. The temperature can be increased during the degradation to enable or hasten the completion of degradation to the ceramic composition 112. Preferably the heating is done from about 400°C to about 600°C, and more preferably to about 500°C. Figures 1 and 2 illustrate the thermal

degradation step for the general scheme (Figure 1) and for specific networked polymers 210 (Figure 2). A fiber-reinforced composite can be made by mixing the precursor 108 with the fibers, then performing the crosslinking to a networked polymer 110, and pyrolysis to a ceramic composition 112.

[0060] Having described the invention, the following examples are given to illustrate specific applications of the invention including specific techniques that can be used to perform the invention. These specific examples are not intended to limit the scope of the invention described in this application. The examples are directed to hydroxy-terminated prepolymers, precursors, and networked polymers where  $n$  is 1, 2, 3, or 4.

#### Example 1

[0061] **Synthesis of 1,4-bis(dimethylaminodimethylsilyl)butadiyne** - A flame dried 250 mL Schlenk flask containing diethyl ether (20 mL) was cooled to  $-78^{\circ}\text{C}$  and  $n$ -butyl lithium (20 mL of 2.4 M in hexane, 48.0 mmol) was added by syringe. After several minutes, hexachlorobutadiene (1.88 mL, 12.0 mmol) was added drop wise via syringe over a 10 minute period. After completion of addition, the cold bath was removed and the mixture stirred at room temperature for 3 hours. The resulting 1,4-dilithio-1,3-butadiyne was used without further purification. The flask was then re-cooled to  $-78^{\circ}\text{C}$  and dimethylaminodimethylchlorosilane (3.6 mL, 24 mmol) was added by syringe. The flask was removed from the cold bath and the reaction mixture stirred at room temperature for 16 hours. At this time,  $^1\text{H}$  NMR analysis indicated complete disappearance of dimethylaminodimethylchlorosilane and formation of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206. Diethyl ether was removed *in vacuo*, and the mixture taken up in a minimum amount of pentane and filtered. Pentane was removed *in vacuo* to give 2.91 g (96%) of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 as a viscous red liquid. IR ( $\text{cm}^{-1}$ ) 2074 (s), ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ).  $^1\text{H}$  NMR (ppm) 2.44 (s), ( $\text{N}(\text{CH}_3)_2$ ), 0.20 (s), ( $\text{Si}(\text{CH}_3)_2$ ).  $^{13}\text{C}$  NMR 87.6, 84.4, ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ) 37.7, ( $\text{N}(\text{CH}_3)_2$ ), -1.2, ( $\text{Si}(\text{CH}_3)_2$ ).

#### Example 2

**[0062] Synthesis of the hydroxy-terminated prepolymer, where  $n=1$**  - To a flame dried 2-necked flask equipped with a reflux condenser was added 1,4-bis(hydroxydimethylsilyl)benzene 200 (1.87 g, 8.25 mmol), 10 mL of THF was injected via syringe followed by 1,4-bis(dimethylsilyl)benzene 202 (0.80 g, 4.12 mmol). The  $(\text{Ph}_3\text{P}_3)\text{RhCl}$  catalyst (0.032 g, 0.032 mmol) was added and the reaction mixture heated to reflux temperature and maintained for 48 hours. IR analysis of the reaction mixture indicated complete disappearance of Si-H stretching band of 1,4-bis(dimethylsilyl)benzene, therefore the reaction was deemed complete. The THF was removed *in vacuo* to give 2.17 g (82 %) of prepolymer 204 as a pale yellow solid. IR ( $\text{cm}^{-1}$ ) 3276 (m, Si-OH), 1072 (s, Si-O).  $^1\text{H}$  NMR (ppm) 7.58 (s), 7.52 (s) ( $\text{C}_6\text{H}_4$ ), 0.37 (s), 0.31 (s) (Si- $\text{CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 140.8, 140.6, 132.4 132.2 ( $\text{C}_6\text{H}_4$ ), 0.88, -0.05 (Si- $\text{CH}_3$ ).

#### Example 3

**[0063] Synthesis of the hydroxy-terminated prepolymer, where  $n=2$**  - To a flame dried 2-necked flask equipped with a reflux condenser was added 1,4-bis(hydroxydimethylsilyl)benzene 200 (2.09 g, 9.25 mmol), 10 mL of THF was injected via syringe followed by 1,4-bis(dimethylsilyl)benzene 202 (1.19 g, 6.16 mmol). The  $(\text{Ph}_3\text{P}_3)\text{RhCl}$  catalyst (0.017 g, 0.018 mmol) was added and the reaction mixture heated to 55-60°C and maintained for 3.5 hours. IR analysis of the reaction mixture indicated complete disappearance of Si-H stretching band of 1,4-bis(dimethylsilyl)benzene, therefore the reaction was deemed complete. The THF was removed *in vacuo* to give 2.79 g (86 %) of prepolymer 204 as a pale yellow solid. IR ( $\text{cm}^{-1}$ ) 3341 (vw, Si-OH), 1078 (s, Si-O).  $^1\text{H}$  NMR (ppm) 7.54 (s), ( $\text{C}_6\text{H}_4$ ), 0.33 (s), (Si- $\text{CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 140.8, 132.2 ( $\text{C}_6\text{H}_4$ ), 0.91 (Si- $\text{CH}_3$ ).

#### Example 4

**[0064] Synthesis of the hydroxy-terminated prepolymer, where  $n=3$**  - To a flame dried 2-necked flask equipped with a reflux condenser was added 1,4-bis(hydroxydimethylsilyl)benzene 200 (0.76 g, 3.36 mmol), 8 mL of THF was injected via syringe followed by 1,4-bis(dimethylsilyl)benzene 202 (0.49 g, 2.52 mmol). The

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(Ph<sub>3</sub>P<sub>3</sub>)RhCl catalyst (0.006 g, 0.007 mmol) was added and the reaction mixture heated to 40°C and maintained for 1.5 hours. IR analysis of the reaction mixture indicated complete disappearance of Si-H stretching band of 1,4-bis(dimethylsilyl)benzene, therefore the reaction was deemed complete. The THF was removed *in vacuo* to give 1.12 g (90 %) of prepolymer 204 as a gray solid. IR (cm<sup>-1</sup>) 3289 (m, Si-OH), 1065 (s, Si-O). <sup>1</sup>H NMR (ppm) 7.55 (s) (C<sub>6</sub>H<sub>4</sub>), 0.33 (s) (Si-CH<sub>3</sub>). <sup>13</sup>C NMR (ppm) 140.8, 132.3 (C<sub>6</sub>H<sub>4</sub>), 0.92 (Si-CH<sub>3</sub>).

## Example 5

**[0065]     Synthesis of the hydroxy-terminated prepolymer, where n=4** - To a flame dried 2-necked flask equipped with a reflux condenser was added 1,4-bis(hydroxydimethylsilyl)benzene 200 (1.34 g, 5.9 mmol), 10 mL of THF was injected via syringe followed by 1,4-bis(dimethylsilyl)benzene 202 (0.92 g, 4.75 mmol). The (Ph<sub>3</sub>P<sub>3</sub>)RhCl catalyst (0.011 g, 0.011 mmol) was added and the reaction mixture heated to 55°C and maintained for 3.5 hours. IR analysis of the reaction mixture indicated incomplete disappearance of Si-H stretching band of 1,4-bis(dimethylsilyl)benzene, therefore the reaction temperature was reduced to 40°C and maintained for 14 hours. At this time IR analysis indicated reaction completion. The THF was removed *in vacuo* to give 1.97 g (88%) of prepolymer 204 as a pale yellow solid. IR (cm<sup>-1</sup>) 3379 (m, Si-OH), 1072 (s, Si-O). <sup>1</sup>H NMR (ppm) 7.62 (s) (C<sub>6</sub>H<sub>4</sub>), 0.41 (s) (Si-CH<sub>3</sub>). <sup>13</sup>C NMR (ppm) 140.8, 132.6 (C<sub>6</sub>H<sub>4</sub>), 0.92 (Si-CH<sub>3</sub>).

## Example 6

**[0066]     Synthesis of the precursor, where n=1** - To a flame dried three-necked flask equipped with a stir bar, reflux condenser, and inlet (syringe needle) and outlet adapters for argon gas was added the previously prepared hydroxy-terminated prepolymer 204 (0.47 g, 0.73 mmol) (see Example 2). A previously prepared sample of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (0.18 g, 0.73 mmol) was dissolved in 10 mL of toluene. A 9.5 mL aliquot of the toluene solution was added to the three-necked flask containing the prepolymer 204. After refluxing the reaction mixture for about 1 hour an additional amount of the toluene solution containing 1,4-

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bis(dimethylaminodimethylsilyl)butadiyne 206 (250  $\mu$ L) was added at time intervals of 15 minutes until the viscosity of the solution had visibly increased and dimethylamine evolution had ceased, as determined by placing a piece of moist litmus paper into the exhaust stream of the argon outlet. After 5 hours an aliquot was removed and a  $^1\text{H}$  NMR spectrum was acquired. After an additional 1 hour at reflux the temperature was reduced to ambient and the solution stirred under a continuous flow of argon overnight. The resulting yellow-brown residue was dissolved in diethyl ether and poured into a separatory funnel containing 20 mL of ice cold  $\text{NH}_4\text{Cl}$  (aq). To break up an emulsion, 10 mL of saturated  $\text{NaCl}$  (aq) was added. The aqueous layer was then extracted twice with 25 mL of diethyl ether. The combined ether extract was dried over anhydrous  $\text{Na}_2\text{SO}_4$  and filtered. The solvent was removed *in vacuo* to give 0.56 g, (95 %) of the precursor 208, where  $n = 2$  as a brown solid. IR ( $\text{cm}^{-1}$ ) 2080 (w,  $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 1072 (s, Si-O).  $^1\text{H}$  NMR (ppm) 7.52 (s) ( $\text{C}_6\text{H}_4$ ), 0.35 (s) 0.31 (s) 0.24 (s) (Si- $\text{CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 141.0, 132.2, ( $\text{C}_6\text{H}_4$ ), 86.9, 84.9 ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 2.07, 1.80, 0.91, 0.59 (Si- $\text{CH}_3$ ).

## Example 7

**[0067]     Synthesis of the precursor, where  $n=2$**  - To a flame dried three-necked flask equipped with a stir bar, reflux condenser, and inlet (syringe needle) and outlet adapters for argon gas was added the previously prepared hydroxy-terminated prepolymer 204 (2.75 g, 2.61 mmol) (see Example 3). A previously prepared sample of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (0.66 g, 2.61 mmol) was dissolved in 10 mL of toluene. A 9.5 mL aliquot of the toluene solution was added to the three-necked flask containing the prepolymer 204. After refluxing the reaction mixture for 1 hour an additional amount of the toluene solution containing 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (250  $\mu$ L) was added at time intervals of 15 minutes until the viscosity of the solution had visibly increased and dimethylamine evolution had ceased, as determined by placing a piece of moist litmus paper into the exhaust stream of the argon outlet. After 5.5 hours an aliquot was removed and a  $^1\text{H}$  NMR spectrum was acquired. An excess amount of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 was detected so an additional small amount of prepolymer 204 was added to the solution. The temperature was reduced to

60°C and the solution stirred under a continuous flow of argon overnight. The resulting yellow-brown residue was dissolved in diethyl ether and poured into a separatory funnel containing 50 mL of ice cold  $\text{NH}_4\text{Cl}$  (aq). To break up an emulsion, 10 mL of saturated  $\text{NaCl}$  (aq) was added. The aqueous layer was then extracted twice with 15 mL of diethyl ether. The combined ether extract was dried over anhydrous  $\text{Na}_2\text{SO}_4$  and filtered. The solvent was removed *in vacuo* to give 2.84 g, (90 %) of precursor 208, where  $n = 4$  as a brown solid. IR ( $\text{cm}^{-1}$ ) 2074 (w,  $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 1078 (s, Si-O).  $^1\text{H}$  NMR (ppm) 7.54 (s) ( $\text{C}_6\text{H}_4$ ), 0.33 (s) (Si- $\text{CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 140.7, 132.3, ( $\text{C}_6\text{H}_4$ ), 86.9, 84.9 ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 1.78, 0.90, (Si- $\text{CH}_3$ ).

#### Example 8

**[0068]    Synthesis of the precursor, where  $n=3$**  - To a flame dried three-necked flask equipped with a stir bar, reflux condenser, and inlet (syringe needle) and outlet adapters for argon gas was added the previously prepared hydroxy-terminated prepolymer 204 (0.59 g, 0.40 mmol) (see Example 4). 8 mL of toluene was added and the solution was stirred until complete dissolution of the prepolymer 204 was achieved. A previously prepared sample of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (0.10 g, 0.40 mmol) was dissolved in 4 mL of toluene and a 3.9 mL aliquot of this solution was added to the three-necked flask containing the prepolymer 204. After refluxing the reaction mixture for 1 to 2 hours an additional amount of the toluene solution containing 1,4-bis(dimethylaminodimethyl- silyl)butadiyne 206 (20-30  $\mu\text{L}$ ) was added at time intervals of 15 minutes until the viscosity of the solution had visibly increased and dimethylamine evolution had ceased, as determined by placing a piece of moist litmus paper into the exhaust stream of the argon outlet. The solution was refluxed for a total of five hours. The temperature was reduced to 45°C and the solution stirred under a continuous flow of argon overnight. The resulting yellow-brown residue was dissolved in diethyl ether and poured into a separatory funnel containing 20 mL of ice cold  $\text{NH}_4\text{Cl}$  (aq). To break up an emulsion, 5 mL of saturated  $\text{NaCl}$  (aq) was added. The aqueous layer was then extracted twice with 15 mL of diethyl ether. The combined ether extract was dried over anhydrous  $\text{Na}_2\text{SO}_4$  and filtered. The solvent was removed *in vacuo* to give 0.64 g, (95 %) of precursor 208, where  $n = 6$  as a brown solid. IR ( $\text{cm}^{-1}$ ) 2080 (w,  $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 1072 (s,



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Si-O).  $^1\text{H}$  NMR (ppm) 7.60 (s) ( $\text{C}_6\text{H}_4$ ), 0.39 (s) ( $\text{Si-CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 140.8, 132.2, ( $\text{C}_6\text{H}_4$ ), 86.9, 84.9 ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 0.92 ( $\text{Si-CH}_3$ ).

## Example 9

**[0069] Synthesis of the precursor, where  $n=4$**  - To a flame dried three-necked flask equipped with a stir bar, reflux condenser, and inlet (syringe needle) and outlet adapters for argon gas was added the previously prepared hydroxy-terminated prepolymer 204 (0.65 g, 0.34 mmol) (see Example 5). A previously prepared sample of 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (0.08 g, 2.61 mmol) was dissolved in 7 mL of toluene. A 6.8 mL aliquot of the toluene solution was added to the three-necked flask containing the prepolymer 204. After refluxing the reaction mixture for 1 hour an additional amount of the toluene solution containing 1,4-bis(dimethylaminodimethylsilyl)butadiyne 206 (20-40  $\mu\text{L}$ ) was added at time intervals of 15 minutes until the viscosity of the solution had visibly increased and dimethylamine evolution had ceased, as determined by placing a piece of moist litmus into the exhaust stream of the argon outlet. The solution was refluxed for a total of 4.5 hours. The temperature was reduced to  $50^\circ\text{C}$  and the solution stirred under a continuous flow of argon overnight. At this time all of the solvent had been removed. Dissolution of the resulting brown solid in toluene followed by a litmus paper test indicated that the solution was neutral. Yield (based upon total amount collected) 0.65 g, (92 %) of precursor 208, where  $n = 8$  as a brown solid. IR ( $\text{cm}^{-1}$ ) 2080 (w,  $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 1059 (s, Si-O).  $^1\text{H}$  NMR (ppm) 7.54 (s) ( $\text{C}_6\text{H}_4$ ), 0.32 (s) 0.31 (s) ( $\text{Si-CH}_3$ ).  $^{13}\text{C}$  NMR (ppm) 140.8, 132.2, ( $\text{C}_6\text{H}_4$ ), 87.0, 84.9 ( $-\text{C}\equiv\text{C}-\text{C}\equiv\text{C}-$ ), 0.92 ( $\text{Si-CH}_3$ ).

## Example 10

**[0070] Crosslinking of precursor, where  $n=1$**  (see Example 6 for synthesis of precursor) - To a platinum thermogravimetric analyzer pan was placed 40.4430 mg of the precursor 208. The sample was then isothermally heated under an atmosphere of dry nitrogen at 150, 200, 350 and  $450^\circ\text{C}$  for 60, 60, 120 and 120 minutes, respectively. After completion of the isothermal curing experiment, the sample was void free and exhibited the characteristics of an elastomeric material; i.e., soft and flexible.

#### Example 11

[0071] **Post crosslinking thermo-oxidative stability study on networked polymer, where  $n=1$**  (see Example 6 for synthesis of precursor and Example 10 for crosslinking of precursor) - Following the crosslinking of the precursor 208, where  $n=1$  the sample was allowed to cool to ambient temperature. The sample was then isothermally heated (in the chamber of a thermogravimetric analyzer) for 120 minutes at 200, 250, 300, and 350°C, respectively, in an air atmosphere. The sample exhibited good oxidative stability over the time frame of the experiment, experiencing an 18.82 % weight loss.

#### Example 12

[0072] **Crosslinking of precursor, where  $n=2$**  (see Example 7 for synthesis of precursor) - To a platinum thermogravimetric analyzer pan was placed 23.8740 mg of the precursor 208. The sample was then isothermally heated under an atmosphere of dry nitrogen at 150, 200, 350 and 450°C for 60, 60, 120 and 120 minutes, respectively. After completion of the isothermal curing experiment the sample was void free and exhibited the characteristics of an elastomeric material; i.e., soft and flexible.

#### Example 13

[0073] **Post crosslinking thermo-oxidative stability study on networked polymer, where  $n=2$**  (see Example 7 for synthesis of precursor and Example 12 for crosslinking of precursor) - Following the crosslinking of the precursor 208, where  $n=2$  the sample was allowed to cool to ambient temperature. The sample was then isothermally heated (in the chamber of a thermogravimetric analyzer) for 120 minutes at 200, 250, 300, and 350°C, respectively, in an air atmosphere. The sample exhibited good oxidative stability over the time frame of the experiment, experiencing an 18.44 % weight loss.

#### Example 14

[0074] **Crosslinking of precursor, where  $n=3$**  (see Example 8 for synthesis of precursor) - To a platinum thermogravimetric analyzer pan was placed 24.6430 mg of the precursor 208, where  $n=3$ . The sample was then isothermally heated under an

atmosphere of dry nitrogen at 150, 200, 350 and 450°C for 60, 60, 120 and 120 minutes, respectively. After completion of the isothermal curing experiment the sample was void free and exhibited the characteristics of an elastomeric material; i.e., soft and flexible.

#### Example 15

**[0075] Post crosslinking thermo-oxidative stability study on networked polymer, where  $n=3$**  (see Example 8 for synthesis of precursor and Example 14 for crosslinking of precursor) - Following the crosslinking of the precursor 208, where  $n=3$  the sample was allowed to cool to ambient temperature. The sample was then isothermally heated (in the chamber of a thermogravimetric analyzer) for 120 minutes at 200, 250, 300, and 350°C, respectively, in an air atmosphere. The sample exhibited good oxidative stability over the time frame of the experiment, experiencing a 13.55 % weight loss.

#### Example 16

**[0076] Crosslinking of precursor, where  $n=4$**  (see Example 9 for synthesis of precursor) - To a platinum thermogravimetric analyzer pan was placed 27.8090 mg of the precursor 208. The sample was then isothermally heated under an atmosphere of dry nitrogen at 150, 200, 350 and 450°C for 60, 60, 120 and 120 minutes, respectively. After completion of the isothermal curing experiment the sample was void free and exhibited the characteristics of an elastomeric material; i.e., soft and flexible.

#### Example 17

**[0077] Post crosslinking thermo-oxidative stability study on networked polymer, where  $n=4$**  (see Example 9 for synthesis of precursor and Example 16 for crosslinking of precursor) - Following the crosslinking of the precursor 208, where  $n=4$  the sample was allowed to cool to ambient temperature. The sample was then isothermally heated (in the chamber of a thermogravimetric analyzer) for 120 minutes at 200, 250, 300, and 350°C, respectively, in an air atmosphere. The sample exhibited good oxidative stability over the time frame of the experiment, experiencing a 19.77 % weight loss.

#### Example 18

**[0078] Determination of the glass transition temperature for networked polymer, where  $n=1$**  - Crosslinked samples were prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. Following the isothermal curing cycle the sample was allowed to cool to ambient temperature. 12.5 mg of the cured sample was weighed into a differential scanning calorimetry pan and the pan was hermetically sealed. The sample was cooled to -70°C and then heated at a rate of 5°C/minute to 400°C. A glass transition was observed at -29°C.

Example 19

**[0079] Determination of the glass transition temperature for networked polymer, where  $n=2$**  - Crosslinked samples were prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. Following the isothermal curing cycle the sample was allowed to cool to ambient temperature. 11.6 mg of the cured sample was weighed into a differential scanning calorimetry pan and the pan was hermetically sealed. The sample was cooled to -70°C and then heated at a rate of 5°C/minute to 400°C. A glass transition was observed at -28°C.

Example 20

**[0080] Determination of the glass transition temperature for networked polymer, where  $n=3$**  - Crosslinked samples were prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. Following the isothermal curing cycle the sample was allowed to cool to ambient temperature. 8.8 mg of the cured sample was weighed into a differential scanning calorimetry pan and the pan was hermetically sealed. The sample was cooled to -70°C and then heated at a rate of 5°C/minute to 100°C. A glass transition was observed at -22°C.

Example 21

**[0081] Determination of the glass transition temperature for networked polymer, where  $n=4$**  - Crosslinked samples were prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. Following the isothermal curing cycle the sample was allowed to cool to ambient temperature. 13.0 mg of the cured sample was weighed into a differential scanning calorimetry pan and the pan was hermetically sealed. The sample was cooled to -70°C and then heated at a rate of 5°C/minute to 100°C. A glass transition was observed at -29°C.

Example 22

**[0082] Post cure dynamic thermogravimetric analysis for networked polymer, where  $n=1$**  - A crosslinked sample was prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. The sample was allowed to cool to ambient temperature and then heated to 1000°C at a rate of 10°C/minute under an atmosphere of dry nitrogen. A one step thermal degradation commenced to occur above 400°C, weight loss was over by 600°C resulting in a char yield of 66.0 %.

Example 23

**[0083] Post cure dynamic thermogravimetric analysis for networked polymer, where  $n=2$**  - A crosslinked sample was prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. The sample was allowed to cool to ambient temperature and then heated to 1000°C at a rate of 10°C/minute under an atmosphere of dry nitrogen. A one step thermal degradation commenced to occur above 400°C, weight loss was over by 600°C resulting in a char yield of 44.0 %.

Example 24

**[0084] Post cure dynamic thermogravimetric analysis for networked polymer, where  $n=3$**  - A crosslinked sample was prepared by heating the precursor 208 for 2 hours

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at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. The sample was allowed to cool to ambient temperature and then heated to 1000°C at a rate of 10°C/minute under an atmosphere of dry nitrogen. A one step thermal degradation commenced to occur above 400°C, weight loss was over by 600°C resulting in a char yield of 35.0 %.

#### Example 25

**[0085] Post cure dynamic thermogravimetric analysis for networked polymer, where  $n=4$**  - A crosslinked sample was prepared by heating the precursor 208 for 2 hours at 150, 250 and 350°C, respectively, in the chamber of a thermogravimetric analyzer. The sample was allowed to cool to ambient temperature and then heated to 1000°C at a rate of 10°C/minute under an atmosphere of dry nitrogen. A one step thermal degradation commenced to occur above 400°C, weight loss was over by 600°C resulting in a char yield of 40.0 %.